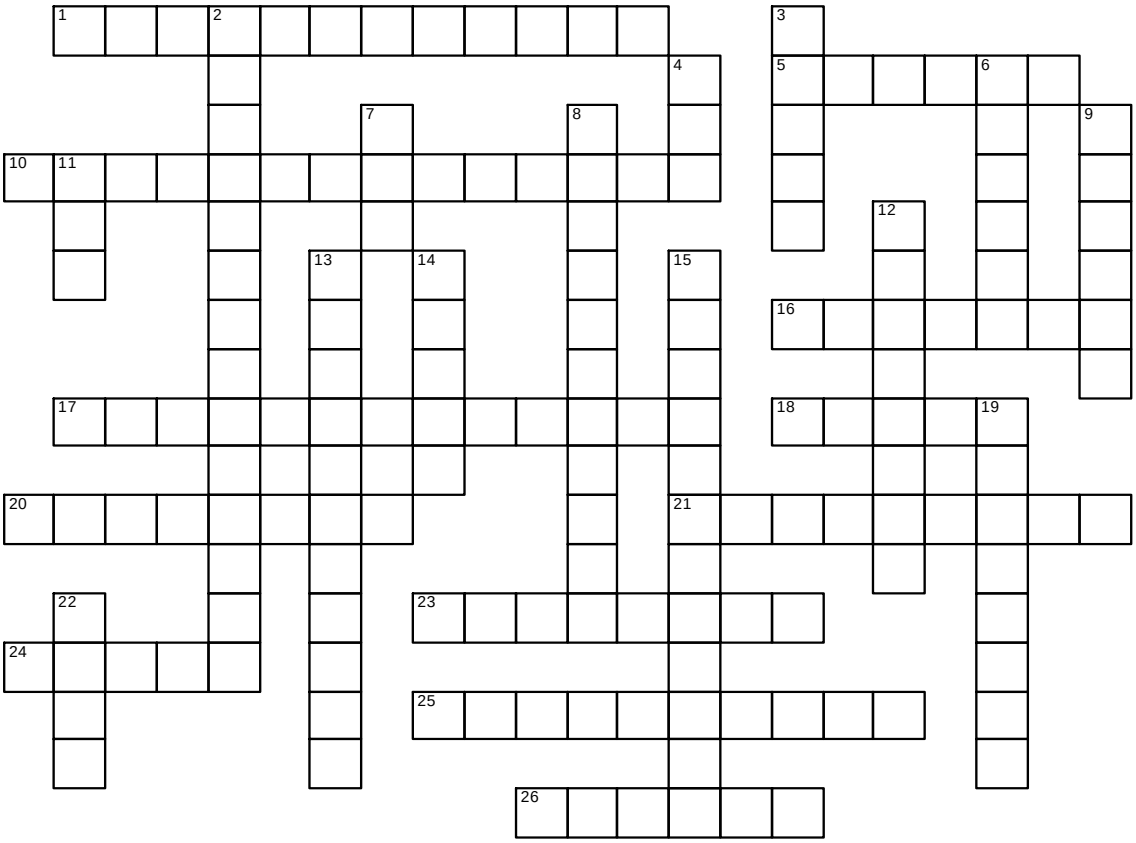
Solubility and Intermolecular Forces

Page References: 262 (fig. 1), 263 (question 5), 264 – 267, 273, 275, 277 - 279



ACROSS

1. The common name for hydrochloric acid
5. The weakest kind of intermolecular force
10. Forces between molecules
16. This is usually the liquid part of a solution
17. Two or more phases
18. Aqua means water in this language
20. The division of matter for which composition can vary (fig. 1 on pg. 262)
21. There are two types of pure substances: elements and ___
23. Two electrons that aren't involved in a bond
24. Polar and ionic solids dissolve in __ solvents
25. Many sites on a molecule for hydrogen bonding usually results in high ___ in water
26. A solution made up of two solids (example)

DOWN

2. Forces within molecules
3. Two metals mixed together
4. A solution made up of gases (example)
6. Liquid solutions can be coloured, but are never ___
7. Water would form London forces of attraction with which molecule(s) in Fig. 6 (pg. 276)
8. A compound that conducts electricity when in an aqueous solution
9. The substance dissolved in a solvent
11. Hydrogen bonding results when hydrogen is bonded to these elements
12. A homogeneous mixture
13. One phase
14. Known as the "universal solvent"
15. The separation of ions when an ionic compound dissolves in water
19. Nonpolar solids dissolve in ___ solvents
22. The number of hydrogen bonds shown in fig 5 (p275)